

Nomenclature Worksheet 5: Ionic Compounds Summary

Name the following compounds:

1. CaF_2 Calcium Fluoride
2. Na_2O Sodium Oxide
3. BaS Barium Sulfide
4. CuSO_4 Copper(II) Sulfate
5. Fe_2O_3 Iron (III) Oxide
6. HgCl_2 Mercury (II) Chloride
7. AgNO_3 Silver Nitrate
8. MgCO_3 Magnesium Carbonate
9. $\text{KC}_2\text{H}_3\text{O}_2$ Potassium Acetate
10. $\text{K}_2\text{Cr}_2\text{O}_7$ Potassium Chromate
11. $\text{Al}(\text{OH})_3$ Aluminum Hydroxide
12. PbBr_2 Lead (II) Bromide
13. ZnSO_3 Zinc Sulfate
14. NaHCO_3 Sodium Bicarbonate
15. NH_4Cl Ammonium Chloride
16. Li_3PO_4 Lithium Phosphate
17. SnCl_2 tin Chloride
18. $\text{Al}(\text{NO}_2)_3$ Aluminum Nitrite
19. Rb_2CrO_4 Rubidium Chromate
20. KMnO_4 Potassium Permanganate
21. CuCl Copper(I) Chloride
22. FeSO_4 Iron Sulfate

Give the formula for each compound:

23. sodium fluoride NaF
24. potassium sulfide K_2S
25. calcium carbonate CaCO_3
26. magnesium hydroxide $\text{Mg}(\text{OH})_2$
27. zinc nitrate $\text{Zn}(\text{NO}_3)_2$
28. silver acetate $\text{AgC}_2\text{H}_3\text{O}_2$
29. copper (II) oxide CuO
30. iron (III) chloride FeCl_3
31. barium chromate BaCrO_4
32. aluminum oxide Al_2O_3
33. lead (II) sulfate $\text{Pb}(\text{SO}_4)$
34. tin (IV) oxalate $\text{Sn}_2(\text{C}_2\text{O}_4)_4$
35. calcium phosphate $\text{Ca}_3(\text{PO}_4)_2$
36. lithium permanganate LiMnO_4
37. mercury (I) nitrate Hg_2NO_3
38. radium sulfite RaSO_3
39. chromium (III) chloride CrCl_3
40. ammonium sulfide $(\text{NH}_4)_2\text{S}$
41. copper (II) acetate $\text{Cu}(\text{C}_2\text{H}_3\text{O}_2)_2$
42. calcium bicarbonate $\text{Ca}(\text{HCO}_3)_2$
43. tin (II) oxide SnO
44. silver sulfite Ag_2SO_3