

## Nomenclature Worksheet 5: Ionic Compounds Summary

Name the following compounds:

1.  $\text{CaF}_2$  \_\_\_\_\_
2.  $\text{Na}_2\text{O}$  \_\_\_\_\_
3.  $\text{BaS}$  \_\_\_\_\_
4.  $\text{CuSO}_4$  \_\_\_\_\_
5.  $\text{Fe}_2\text{O}_3$  \_\_\_\_\_
6.  $\text{HgCl}_2$  \_\_\_\_\_
7.  $\text{AgNO}_3$  \_\_\_\_\_
8.  $\text{MgCO}_3$  \_\_\_\_\_
9.  $\text{KC}_2\text{H}_3\text{O}_2$  \_\_\_\_\_
10.  $\text{K}_2\text{Cr}_2\text{O}_7$  \_\_\_\_\_
11.  $\text{Al}(\text{OH})_3$  \_\_\_\_\_
12.  $\text{PbBr}_2$  \_\_\_\_\_
13.  $\text{ZnSO}_3$  \_\_\_\_\_
14.  $\text{NaHCO}_3$  \_\_\_\_\_
15.  $\text{NH}_4\text{Cl}$  \_\_\_\_\_
16.  $\text{Li}_3\text{PO}_4$  \_\_\_\_\_
17.  $\text{SnCl}_2$  \_\_\_\_\_
18.  $\text{Al}(\text{NO}_2)_3$  \_\_\_\_\_
19.  $\text{Rb}_2\text{CrO}_4$  \_\_\_\_\_
20.  $\text{KMnO}_4$  \_\_\_\_\_
21.  $\text{CuCl}$  \_\_\_\_\_
22.  $\text{FeSO}_4$  \_\_\_\_\_

Give the formula for each compound:

23. sodium fluoride \_\_\_\_\_
24. potassium sulfide \_\_\_\_\_
25. calcium carbonate \_\_\_\_\_
26. magnesium hydroxide \_\_\_\_\_
27. zinc nitrate \_\_\_\_\_
28. silver acetate \_\_\_\_\_
29. copper (II) oxide \_\_\_\_\_
30. iron (III) chloride \_\_\_\_\_
31. barium chromate \_\_\_\_\_
32. aluminum oxide \_\_\_\_\_
33. lead (II) sulfate \_\_\_\_\_
34. tin (IV) oxalate \_\_\_\_\_
35. calcium phosphate \_\_\_\_\_
36. lithium permanganate \_\_\_\_\_
37. mercury (I) nitrate \_\_\_\_\_
38. radium sulfite \_\_\_\_\_
39. chromium (III) chloride \_\_\_\_\_
40. ammonium sulfide \_\_\_\_\_
41. copper (II) acetate \_\_\_\_\_
42. calcium bicarbonate \_\_\_\_\_
43. tin (II) oxide \_\_\_\_\_
44. silver sulfite \_\_\_\_\_