

Nomenclature Worksheet 4:  
Ionic Compounds Containing Transition Metals

Please complete the following table:

Name of Ionic Compound	Formula of Ionic Compound
1. Copper (II) sulfate	$CuSO_4$
2. Copper (I) oxide	$Cu_2O$
3. Chromium (III) cyanide	$Cr(CN)_3$
4. Cobalt (II) hydroxide	$Co(OH)_2$
5. Silver bromide	$AgBr$
6. Zinc nitrate	$Zn(NO_3)_2$
7. Iron (III) acetate	$Fe(C_2H_3O_2)_3$
8. Lead (IV) sulfate	$Pb_2(SO_4)_4$
9. Iron (II) Chloride	$FeCl_2$
10. Lead Sulfate	$PbSO_3$
11. Calcium Carbonate	$Ca_2(CO_3)_3$ $Ca_2(CO_3)_3$
12. Silver Nitrate	$AgNO_3$
13. Zinc Cyanide	$Zn(CN)_2$
14. Copper (I) Chlorate	$CuClO_3$
15. Chromium (III) Hydroxide	$Cr(OH)_3$
16. Mercury (I) Oxide	$Hg_2O$