

Nomenclature Worksheet 3: Ionic Compounds Containing Polyatomic Ions

Please complete the following table:

Name of Ionic Compound	Formula of Ionic Compound
1. Sodium chromate ⁺¹ ₋₂	Na_2CrO_4
2. Calcium carbonate ⁺² ₋₂	CaCO_3
3. Magnesium nitrate ⁺² ₋₁	$\text{Mg}(\text{NO}_3)_2$
4. Aluminum sulfate ⁺³ ₋₂	$\text{Al}_2(\text{SO}_4)_3$
5. Lithium phosphate ⁺¹ ₋₃	$\text{Li}_3(\text{PO}_4)$
6. Ammonium chloride ⁺¹ ₋₁	NH_4Cl
7. Cesium chlorate ⁺¹ ₋₁	CsClO_3
8. Potassium sulfate ⁺¹ ₋₂	K_2SO_4
9. Barium acetate ⁺² ₋₁	$\text{Ba}(\text{C}_2\text{H}_3\text{O}_2)_2$
10. Rubidium cyanide ⁺¹ ₋₁	RbCN
11. Potassium Acetate	KCH_3CO_2
12. Magnesium Phosphate	$\text{Mg}_3(\text{PO}_4)_2$
13. Aluminium Chlorate	$\text{Al}(\text{ClO}_3)_3$
14. Calcium Sulfate	CaSO_4
15. Strontium Bicarbonate	$\text{Sr}(\text{HCO}_3)_2$
16. Sodium Nitrate	NaNO_3
17. Lithium Carbonate	Li_2CO_3
18. Barium Nitrate	$\text{Ba}(\text{NO}_3)_2$
19. Cesium Chromate	Cs_2CrO_4
20. Ammonium Hydroxide	NH_4OH