

CHEMISTRY SINGLE REPLACEMENT REACTION WORKSHEET

Using the Activity Series Table, complete the following reactions by writing the products that are formed. Be sure to Balance each equation. If NO single replacement reaction occurs, write *NR* to the right of the arrow.

1. $\text{Ag} + \text{KNO}_3 \rightarrow$
2. $\text{Zn} + \text{AgNO}_3 \rightarrow$
3. $\text{Al} + \text{H}_2\text{SO}_4 \rightarrow$
4. $\text{Cl}_2 + \text{KI} \rightarrow$
5. $\text{Li} + \text{H}_2\text{O} \rightarrow$
6. $\text{Cu} + \text{FeSO}_4 \rightarrow$
7. $\text{Na} + \text{H}_2\text{O} \rightarrow$
8. $\text{Fe} + \text{Pb}(\text{NO}_3)_2 \rightarrow$
9. $\text{Cu} + \text{H}_2\text{O} \rightarrow$
10. $\text{Cu} + \text{Al}_2(\text{SO}_4)_3 \rightarrow$
11. $\text{Al} + \text{Pb}(\text{NO}_3)_2 \rightarrow$
12. $\text{Cl}_2 + \text{NaI} \rightarrow$
13. $\text{Fe} + \text{AgCH}_3\text{COO} \rightarrow$
14. $\text{Al} + \text{CuCl}_2 \rightarrow$
15. $\text{Br}_2 + \text{CaI}_2 \rightarrow$
16. $\text{Al} + \text{HCl} \rightarrow$
17. $\text{Mg} + \text{HCl} \rightarrow$
18. $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow$
19. $\text{Fe} + \text{CuSO}_4 \rightarrow$
20. $\text{Cl}_2 + \text{MgI}_2 \rightarrow$