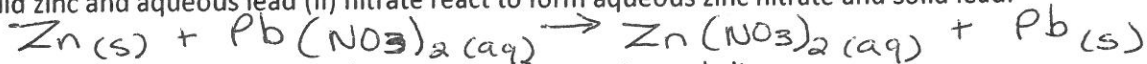


Name Key Study Guide "Chemical Reactions" Date \_\_\_\_\_ Period \_\_\_\_\_

Part I: Write the skeleton equation (don't balance) for the following word equation. (Remember to include catalyst and heat added when used, and phases.)

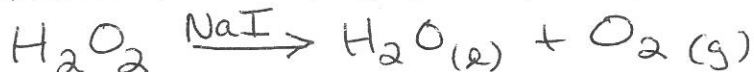
1. Solid zinc and aqueous lead (II) nitrate react to form aqueous zinc nitrate and solid lead.



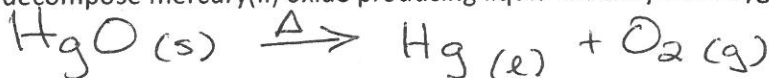
2. Hydrogen and nitrogen monoxide gas react to form water and nitrogen.



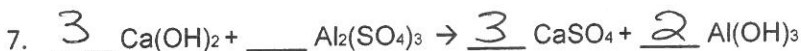
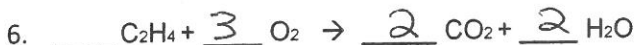
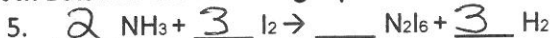
3. Sodium iodide is added to speed up the decomposition of hydrogen peroxide producing water vapor and oxygen.



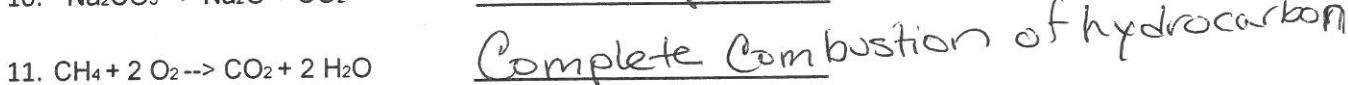
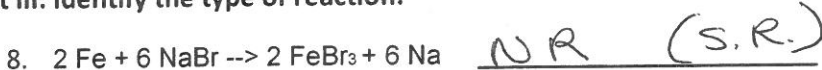
4. Heat is added to decompose mercury(II) oxide producing liquid mercury and oxygen.



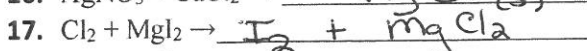
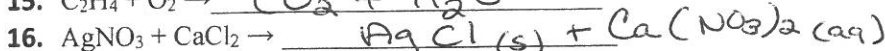
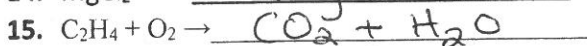
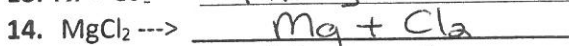
Part II. Balance the following equations.



Part III: Identify the type of reaction:



Part IV: Predict the products:



19. In order for a double displacement reaction to occur, one or more of the following products need to be produced: solid precipitate, (H<sub>2</sub>O) molecular compd. and gas.

Part V: Throw it all together now. Predict the products and balance:

20. Solid zinc reacts with hydrochloric acid (aq) to produce .....

