

The Atoms Family
Atomic Math Challenge

Name ANSWER KEY

8
O
Oxygen
15.999

Atomic #
Symbol
NAME
Atomic Mass

Atomic number equals the number of
PROTONS or ELECTRONS
mass number (IF NEUTRALLY CHARGED)
Atomic mass equals the number of
PROTONS + NEUTRONS

↳ Rounded Atomic Mass to get Mass Number

8
O
<u>Oxygen</u>
15.999

Atomic # = 8
Atomic Mass = 16
of Protons = 8
of Neutrons = 8
of Electrons = 8

30
<u>Zn</u>
Zinc
65.39

Atomic # = 30
Atomic Mass = 65
of Protons = 30
of Neutrons = 35
of Electrons = 30

3
<u>Li</u>
Lithium
6.941

Atomic # = 3
Atomic Mass = 7
of Protons = 3
of Neutrons = 4
of Electrons = 3

14
<u>Si</u>
Silicon
28.086

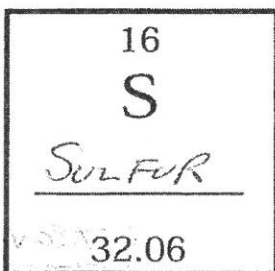
Atomic # = 14
Atomic Mass = 28
of Protons = 14
of Neutrons = 14
of Electrons = 14

5
<u>B</u>
Boron
10.81

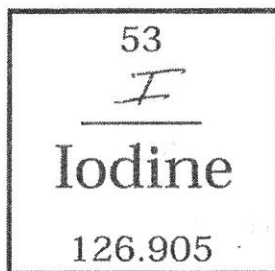
Atomic # = 5
Atomic Mass = 11
of Protons = 5
of Neutrons = 6
of Electrons = 5

35
<u>Br</u>
Bromine
79.904

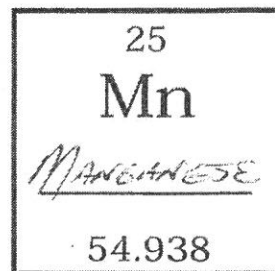
Atomic # = 35
Atomic Mass = 80
of Protons = 35
of Neutrons = 45
of Electrons = 35



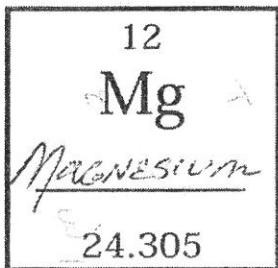
Atomic # = $\frac{16}{\#}$
~~Atomic Mass~~ # = $\frac{32}{\#}$
 # of Protons = $\frac{16}{\#}$
 # of Neutrons = $\frac{16}{\#}$
 # of Electrons = $\frac{16}{\#}$



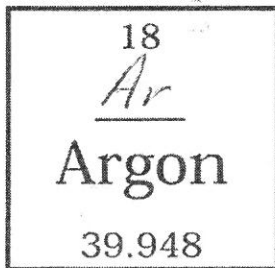
Atomic # = $\frac{53}{\#}$
~~Atomic Mass~~ # = $\frac{127}{\#}$
 # of Protons = $\frac{53}{\#}$
 # of Neutrons = $\frac{74}{\#}$
 # of Electrons = $\frac{53}{\#}$



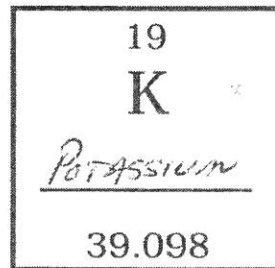
Atomic # = $\frac{25}{\#}$
~~Atomic Mass~~ # = $\frac{55}{\#}$
 # of Protons = $\frac{25}{\#}$
 # of Neutrons = $\frac{30}{\#}$
 # of Electrons = $\frac{25}{\#}$



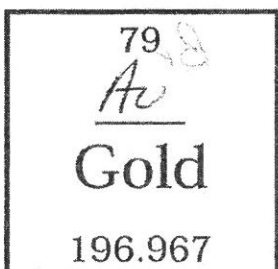
Atomic # = $\frac{12}{\#}$
~~Atomic Mass~~ # = $\frac{24}{\#}$
 # of Protons = $\frac{12}{\#}$
 # of Neutrons = $\frac{12}{\#}$
 # of Electrons = $\frac{12}{\#}$



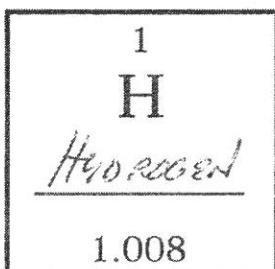
Atomic # = $\frac{18}{\#}$
~~Atomic Mass~~ # = $\frac{40}{\#}$
 # of Protons = $\frac{18}{\#}$
 # of Neutrons = $\frac{22}{\#}$
 # of Electrons = $\frac{18}{\#}$



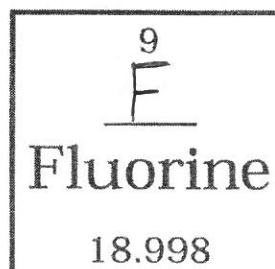
Atomic # = $\frac{19}{\#}$
~~Atomic Mass~~ # = $\frac{39}{\#}$
 # of Protons = $\frac{19}{\#}$
 # of Neutrons = $\frac{20}{\#}$
 # of Electrons = $\frac{19}{\#}$



Atomic # = $\frac{79}{\#}$
~~Atomic Mass~~ # = $\frac{197}{\#}$
 # of Protons = $\frac{79}{\#}$
 # of Neutrons = $\frac{118}{\#}$
 # of Electrons = $\frac{79}{\#}$



Atomic # = $\frac{1}{\#}$
~~Atomic Mass~~ # = $\frac{1}{\#}$
 # of Protons = $\frac{1}{\#}$
 # of Neutrons = $\frac{0}{\#}$
 # of Electrons = $\frac{1}{\#}$



Atomic # = $\frac{9}{\#}$
~~Atomic Mass~~ # = $\frac{19}{\#}$
 # of Protons = $\frac{9}{\#}$
 # of Neutrons = $\frac{10}{\#}$
 # of Electrons = $\frac{9}{\#}$